

1. Draw Lewis structures for the following **covalent** compounds:

H <sub>2</sub>	F <sub>2</sub>	O <sub>2</sub>
N <sub>2</sub>	HCl	HBr
H <sub>2</sub> O	H <sub>2</sub> S	NH <sub>3</sub>
CH <sub>4</sub>	CCl <sub>4</sub>	NCl <sub>3</sub>
CS <sub>2</sub>	CHCl <sub>3</sub>	HCN

2. Draw Lewis structures for the following **covalent** compounds:

$\text{H}_2\text{O}_2$	$\text{CH}_2\text{O}$ (carbon is central)
$\text{C}_2\text{H}_6$	$\text{C}_2\text{H}_4$
$\text{C}_2\text{H}_2$	$\text{CO}_2$ (carbon is central)
$\text{C}_2\text{H}_5\text{Cl}$	$\text{C}_2\text{H}_5\text{OH}$
$\text{CH}_3\text{OCH}_3$	$\text{CH}_3\text{NH}_2$